

[21-BS 224]

AT THE END OF SECOND SEMESTER, (CBCS PATTERN)

CHEMISTRY-II - ORGANIC AND GENERAL CHEMISTRY

(COMMON FOR B.Sc (CHEMISTRY) and B.Voc Courses)

(UG PROGRAM (4 YEARS HONORS))

(w.e.f. Admitted batch 2020-21)

Time: 3 Hours

Max. Marks: 75

SECTION A (5 × 5 = 25 marks)

Answer any FIVE of the following questions

1. Write the conformations of cyclohexane.
సైక్లోహెక్సేన్ యొక్క కన్ఫర్మేషన్లను వ్రాయండి.
2. Discuss reactivity in alkanes.
ఆల్కేన్ల చర్యాత్మకతను చర్చించండి.
3. Write mechanism of addition of HX to an unsymmetrical alkene.
అసౌష్ఠవ ఆల్కేన్లతో HX జరుపు సంకలన చర్యా సంవిధానాన్ని వ్రాయండి.
4. Discuss hydration of alkynes.
ఆల్కైన్లల హైడ్రేషన్ చర్యను తెలపండి.
5. Define Ring activating and deactivating groups.
వలయ ఉత్తేజిత మరియు నిరుత్తేజిత సమూహాలను నిర్వచించండి.
6. Apply VB theory to $Ni(CO)_4$ molecule.
 $Ni(CO)_4$ అణువుకు VB సిద్ధాంతాన్ని అనువర్తించండి.
7. Describe HSAB principle and its importance.
HSAB నియమం మరియు దాని ప్రాముఖ్యతను చర్చించండి.
8. Define Racemic mixtures.
రేసిమిక్ మిశ్రమాన్ని నిర్వచించండి.

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SECTION B — (5 × 10 = 50 marks)

Answer ALL questions

9. (a) Write the preparation of alkanes and explain the mechanism of halogenation of alkanes by free radicals substitution mechanism.

(b) Explain stability of cycloalkanes by Baeyer strain theory with limitations.

10. (a) Write the preparation of alkenes and discuss Saytzeff and Hofmann eliminations in alkenes.

(b) Discuss the addition of HX to a conjugated diene by 1,2 and 1,4-addition mechanism.

11. (a) Define Huckel's rule and discuss aromaticity in Benzoid and Non-Benzoid compound with examples.

(b) Explain mechanism of following electrophilic substitution reactions.

(i) Nitration of benzene

(ii) Friedel-Crafts alkylation of benzene

(iii) Benzene ring - Electrophilic aromatic substitution

12. (a) Define chemical adsorption and derive an equation for Langmuir adsorption isotherm.

(b) Define LCAO method and construct M.O diagram for O2 molecule.

13. (a) Define optical isomerism and explain optical isomerism in lactic acid and Tartaric acid.

(b) Discuss D, L and E, Z configuration in carbon compounds with examples.

(c) Discuss the configuration of D, L and E, Z in carbon compounds with examples.

(d) Discuss the configuration of D, L and E, Z in carbon compounds with examples.

(e) Discuss the configuration of D, L and E, Z in carbon compounds with examples.

(f) Discuss the configuration of D, L and E, Z in carbon compounds with examples.

(g) Discuss the configuration of D, L and E, Z in carbon compounds with examples.

(h) Discuss the configuration of D, L and E, Z in carbon compounds with examples.

(i) Discuss the configuration of D, L and E, Z in carbon compounds with examples.

(j) Discuss the configuration of D, L and E, Z in carbon compounds with examples.

(k) Discuss the configuration of D, L and E, Z in carbon compounds with examples.

(l) Discuss the configuration of D, L and E, Z in carbon compounds with examples.

(m) Discuss the configuration of D, L and E, Z in carbon compounds with examples.

(n) Discuss the configuration of D, L and E, Z in carbon compounds with examples.

(o) Discuss the configuration of D, L and E, Z in carbon compounds with examples.

(p) Discuss the configuration of D, L and E, Z in carbon compounds with examples.

(q) Discuss the configuration of D, L and E, Z in carbon compounds with examples.

(r) Discuss the configuration of D, L and E, Z in carbon compounds with examples.

(s) Discuss the configuration of D, L and E, Z in carbon compounds with examples.

(t) Discuss the configuration of D, L and E, Z in carbon compounds with examples.

(u) Discuss the configuration of D, L and E, Z in carbon compounds with examples.

(v) Discuss the configuration of D, L and E, Z in carbon compounds with examples.

(w) Discuss the configuration of D, L and E, Z in carbon compounds with examples.

(x) Discuss the configuration of D, L and E, Z in carbon compounds with examples.